Skill Practice 63

*More EQ Practice*

Name: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Date: \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

Hour: \_\_\_\_\_

1. When 0.100 mol H2S gas was put into a 10.0 L vessel and heated to 1132oC, it gave an equilibrium mixture containing 0.0285 mol H2 gas along with some S2 gas.
2. Calculate Kc at this temperature.
3. Calculate Kp at this temperature.
4. If Kc for the following reaction is 4.58x10-4, what is Kp at 420oC?

SO2Cl2 (g) 🡨🡪 SO2 (g) + Cl2 (g)

1. Given the information in question 2 above, consider the following. a 2.5 L reaction vessel was filled with 0.84 mol of SO2Cl2. After the reaction comes to equilibrium, what is the concentration of SO2Cl2, SO2 and Cl2?
2. Consider the following chemical equation: 2 N2O5 🡨🡪 2 N2 + 5 O2. If Kc = 0.345 at a temperature of 245 K, what is Kp at this temperature?